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Chohey (deceased) was Editor-in-Chief at Chemical Engineering magazine, with more than 45 years' experience in engineering and publishing. Handbook of Chemical Engineering Calculations, Fourth ... Handbook of Chemical Engineering

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 Assume gas is at standard conditions, that  
 is,  $1 \text{ g} \cdot \text{mol}$  gas at  $1 \text{ atm}$  ( $101.3 \text{ kPa}$ )  
 pressure and  $0 \text{ C}$  ( $273 \text{ K}$ , or  $492 \text{ R}$ ),

occupying a volume of  $22.4 \text{ L}$ . 2. Compute  
 the gas constant. Apply suitable  
 conversion factors and obtain the gas  
 constant in various units. Use  $PV = RT$ ;  
 that is,  $R = PV/T$ . Thus, a. SECTION 1  
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 Chilton—Chemical Engineers' Handbook,  
 McGraw-Hill, 1973.) 1.9 HEAT CAPACITY OF  
 REAL GASES Calculate the heat capacity  $C_p$   
 of ethane vapor at  $400 \text{ K}$  ( $260 \text{ F}$ ) and  $50$   
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Calculation Procedure 1. Assume a basis. Assume gas is at standard conditions, that is, 1 g·mol gas at 1 atm (101.3 kPa) pressure and 0 C (273 K, or 492 R), occupying a volume of 22.4 L. 2. Compute the gas constant. Apply suitable conversion factors and obtain the gas

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Handbook of Chemical Engineering Calculations Nicholas Chohey \* Provides detailed procedures for performing hundreds of chemical engineering calculations along with fully worked-out examples

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Hicks holds a bachelor's degree in Mechanical Engineering from Cooper Union School of Engineering in New York. He is the author or coauthor of more than 20 books in engineering and related fields, including Civil Engineering Formulas, Handbook of Mechanical Engineering Calculations, Handbook of Energy

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Avinash Gupta, Ph.D. Senior Principal Chemical Engineer Chevron Lummus Global Bloomfield, NJ

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