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Naming Inorganic Compounds the Easy Way! **Inorganic Nomenclature** Naming Ionic and Molecular Compounds | How to Pass Chemistry

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Naming Ionic Compounds

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nomenclature of organic chemistry.IUPAC

nomenclature of inorganic chemistry -

WikipediaAnswer. 1.C; Calcium +

Carbonate --> $Ca^{2+} + CO_3^{2-} \rightarrow CaCO_3$.

2.D; $FeO \rightarrow Fe + O_2 \rightarrow$ Iron must have

a charge of +2 to make a neutral

compound --> $Fe^{2+} + O_2 \rightarrow$ Iron (II)

Oxide. 3.A; $Al(NO_3)_3 \rightarrow Al^{3+} + (NO_3)^{-3}$ -

-> Aluminum nitrate. 4.B; Phosphorus

trichloride --> $P + 3Cl \rightarrow PCl_3$. 5.D, $LiClO$

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of chemical nomenclature can be found in

Principles of Chemical Nomenclature.

Greater detail can be found in the

Nomenclature of Inorganic Chemistry,

colloquially known as the Red Book , and

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—covalent compounds that contain only

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procedure similar to that used for simple

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as needed to indicate the number of

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the following steps:2.8: Naming Inorganic

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Book, is a collection of recommendations

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chemistry. Part II.2. Nomenclature of

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anions and ligands, Pure Appl. Chem.,

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simply refers to a naming convention, the

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chemistry also has an IUPAC

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In chemical nomenclature, the IUPAC

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Answer. 1.C; Calcium + Carbonate --> $\text{Ca}^{2+} + \text{CO}_3^{2-} \rightarrow \text{CaCO}_3$. 2.D; $\text{FeO} \rightarrow \text{Fe} + \text{O}^{2-} \rightarrow$ Iron must have a charge of +2 to make a neutral compound --> $\text{Fe}^{2+} + \text{O}^{2-} \rightarrow$ Iron (II) Oxide. 3.A; $\text{Al}(\text{NO}_3)_3 \rightarrow \text{Al}^{3+} + (\text{NO}_3)^{-3} \rightarrow$ Aluminum nitrate. 4.B; Phosphorus trichloride --> $\text{P} + 3\text{Cl} \rightarrow \text{PCl}_3$. 5.D, LiClO ...

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An overall summary of chemical nomenclature can be found in Principles of Chemical Nomenclature. Greater detail can be found in the Nomenclature of Inorganic Chemistry, colloquially known as the Red Book, and in the related publications for organic compounds (the Blue Book) and polymers (the Purple Book). It should be noted that many compounds may have non-systematic or semi-systematic names (some of which are not accepted by IUPAC for several reasons, for example because they are ...

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[Nomenclature of Inorganic Chemistry: IUPAC Recommendations 2005](#) is the definitive guide for scientists working in academia or industry, for scientific publishers of books, journals and databases, and for organisations requiring internationally approved nomenclature in a legal or regulatory environment. [Nomenclature Of Inorganic Chemistry Inorganic](#)

The Rules of Inorganic Nomenclature (the 'Red Book'), first published in 1958 by the International Union of Pure and Applied Chemistry (IUPAC), was most recently updated as Nomenclature of Inorganic Chemistry 1990. This new edition of the Red Book clarifies and updates recommendations concerning the names and formulae of inorganic compounds and reflects major recent developments in inorganic chemistry.

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Binary Inorganic Compounds Binary covalent compounds —covalent compounds that contain only two

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Nomenclature simply refers to a naming convention, the International Union of Pure and Applied Chemistry (IUPAC) naming conventions referred to as IUPAC nomenclature is used to systematically name various chemical compounds. Topic: Properties of Matter

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The Rules of Inorganic Nomenclature first published in 1958 was most recently updated as Nomenclature of Inorganic Chemistry 1990. The present book supersedes not only Red Book I but also, where appropriate, Nomenclature of Inorganic Chemistry II (Red Book II). Nomenclature of inorganic chemistry. Part II.2. Nomenclature of hydrides of nitrogen and derived cations, anions and ligands, Pure Appl. Chem., 1982, 54, 2545-2552. A PDF of the printed version is available.

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