

## Chapter 4 Atomic Structure Section 4 1 Studying Atoms

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 New System of Chemical Philosophy, John Dalton presented his  
 symbols for the elements known at that time and their possible  
 combinations. John Dalton Although the concept of the atom was  
 revived in the eighteenth century, it took another hundred years  
 before significant Chapter 4: The Structure of the Atom Chapter 4:  
 The Structure of the Atom. 86 Chapter 4. What You'll Learn. You  
 will identify the experi- ments that led to the devel- opment of the  
 nuclear model of atomic structure. You will describe the struc-  
 ture of the atom and differ- entiate among the subatomic  
 particles that comprise it. Chapter 4: The Structure of the  
 Atom Section 4.2 Structure of the Nuclear Atom One change to  
 Dalton's atomic theory is that atoms are divisible into subatomic

particles: Electrons, protons, and neutrons are examples of these  
 fundamental particles There are many other types of particles,  
 but we will study these three Chapter 4 Atomic Structure - Henry  
 County School District Section 4.1 Defining the Atom. The Greek  
 philosopher Democritus (460 B.C. - 370 B.C.) was among the first  
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 "atomos") He believed that atoms were . indivisible. and .  
 indestructible. His ideas did agree with later scientific theory, but  
 did not explain chemical behavior, and was Chapter 4 Atomic  
 Structure - cville.k12.ky.us Section 4.1 - Studying Atoms  
 Democritus believed that all matter consisted of extremely small  
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 of properties. Chapter 4 Atomic Structure -  
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 indivisible particles called atoms. 2. Atoms of the same element  
 are identical. The atom of any one element are different from  
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 Structure Flashcards | Quizlet Section 4.1 Defining the Atom  
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 is always being the best friend for spending little time in. Section  
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 Atoms. OBJECTIVES: Explain what makes . elements. and .  
 isotopes. different from each other. Calculate the number of  
 neutrons in an atom. Calculate the atomic mass of an element.  
 Explain why chemists use the periodic table. Chapter 4 "Atomic  
 Structure" Section 4.2 Structure of the Nuclear Atom OBJECTIVES:  
 Describe the structure of atoms, according to the Rutherford  
 atomic model. 10. Section 4.2 Structure of the Nuclear Atom One  
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Structure Section 4.1 Studying Atoms Chapter 4.2 Structure of the Nuclear Atom Subatomic Particles What are three kinds of subatomic particles? Much of Dalton's atomic theory is accepted today. • One important change, however, is that atoms are now known to be divisible. • They can be broken down into even smaller, more fundamental particles, called subatomic particles. Section 4.1 Defining the Atom. The Greek philosopher Democritus (460 B.C. – 370 B.C.) was among the first to suggest the existence of atoms (from the Greek word "atomos") He believed that atoms were . indivisible. and . indestructible. His ideas did agree with later scientific theory, but did not explain chemical behavior, and was

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Section 4.3 Distinguishing Among Atoms. OBJECTIVES: Explain what makes . elements. and . isotopes. different from each other. Calculate the number of neutrons in an atom. Calculate the atomic mass of an element. Explain why chemists use the periodic table.

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Section 4.2 Structure of the Nuclear Atom One change to Dalton's atomic theory is that atoms are divisible into subatomic particles: Electrons, protons, and neutrons are examples of these fundamental particles There are many other types of particles, but we will study these three

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Section 4.2 Structure of the Nuclear Atom OBJECTIVES: Describe the structure of atoms, according to the Rutherford atomic model. 10. Section 4.2 Structure of the Nuclear Atom One change to Dalton's atomic theory is that atoms are divisible into subatomic particles: Electrons Protons Neutrons. 11.

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View full document. Chapter 4 Notes – Atomic Structure Section 4.2 – Properties of Subatomic Particles Protons, electrons, and neutrons are subatomic particles. Protons – a positively charged subatomic particle that is found in the nucleus of an atom. - Every nucleus must contain at least one proton.

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Section 4.1 – Studying Atoms Democritus believed that all matter consisted of extremely small particles that could not be divided. He called these particles atoms from the Greek word "atomos", which meant indivisible. He thought that there were different types of atoms with specific sets of properties.

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All elements are composed of tiny indivisible particles called atoms. 2. Atoms of the same element are identical. The atom of any one element are different from those of any other element. 3. Atoms of different elements can physically mix together or can chemically combine in simple whole number ratios to form compounds. 4.

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