

Introduction To Atmospheric Chemistry Solutions Manual

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 INTRODUCTION TO ATMOSPHERIC CHEMISTRY
 Introduction to Atmospheric Chemistry: A long-term Perspective
 Solutions for Selected Exercises

*Introduction To Atmospheric
 Chemistry Solutions Manual*

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 1 MEASURES OF ATMOSPHERIC COMPOSITION
 1. 1.1 MIXING RATIO 1. 1.2 NUMBER DENSITY 2. 1.3 PARTIAL PRESSURE
 6. PROBLEMS 10. 1.1 Fog formation 10. 1.2 Phase partitioning of water in cloud
 10. 1.3 The ozone layer 10. 2 ATMOSPHERIC PRESSURE 12. 2.1 MEASURING ATMOSPHERIC PRESSURE
 12. 2.2 MASS OF THE ATMOSPHERE 13. 2.3 VERTICAL PROFILES OF PRESSURE AND TEMPERATURE
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 ...Introduction to Atmospheric Chemistry (Princeton University Press, 1999). They are arranged following the different chapters of the book. In recent years I have added to my course lectures a chapter 14, 'Aerosol Chemistry' and a chapter 15, 'Mercury in the Environment'. I have included here problems to support these chapters.
 INTRODUCTION TO ATMOSPHERIC CHEMISTRY Daniel J. Jacob is the Gordon McKay Professor of Atmospheric Chemistry and Environmental Engineering at Harvard University. He has taught the undergraduate atmospheric chemistry course at Harvard since 1992. He has published over 100 research papers in atmospheric chemistry journals.
 Introduction to Atmospheric Chemistry: Daniel J. Jacob ...
 Importance of Atmospheric Chemistry • Atmosphere is very thin and fragile! - Earth diameter = 12,740 km - Earth mass ~ 6×10^{24} kg - Atmospheric mass ~ 5.1×10^{18} kg - 99% of atmospheric mass below ~ 50 km - Solve in class: order of magnitude of mass of the oceans? Mass of entire human population?
 Lecture 1: Introduction to Atmospheric Chemistry
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 1. Introduction to Environmental Chemistry Readings: Chapter 1 Course introductory comments Review of Important Chemical Principles and Units of Measure (handout) Introductory Lecture Definitions Box Models Atmospheric Residence Times
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 Introduction to Atmospheric Chemistry is a concise, clear review of the fundamental aspects of atmospheric chemistry. In ten succinct chapters, it reviews our basic understanding of the chemistry of the Earth's atmosphere and discusses current environmental issues, including air pollution, acid rain, the ozone hole, and global change.
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 Solution For the air in balloon at ground level: $p_i = 1013$ hPa $T_i = 20$ C = 293 K For the air in balloon at 900 hPa $p_f = 900$ hPa $T_f = 87.43$ C (from solution to Exercise 3.22) = 360.43 K Suppose the air in the balloon goes from its initial to its final state in two steps: 1
 Solutions for Selected Exercises
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 40 Consider the typical vertical profile of ozone number density measured over the US shown in the opposite figure.
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 Guy Brasseur Max Planck Institute for Meteorology, Hamburg, Germany and National Center for Atmospheric

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